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atoms of different elements have different masses. (all. oxygen atoms have the same mass; and the mass of oxygen. is always different from every other element's mass)

□ Compounds contain atoms of more than

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one element. (ex: NaCl) □ In a particular compound, atoms of different elements always combine in the same way.

Chapter 4: Atomic Structure Notes
Flashcards | Quizlet

Name: _____ Period: _____ Unit 2

Atomic Structure- Guided Notes The

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Periodic Table _____ is credited for creating the first version of the modern periodic table _____ is a chart showing all the elements arranged in columns in such a way that all the elements in a given column exhibit similar chemical properties
The periodic table is organized in order of _____ The periodic is organized in&mldr

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most often credited with the discovery of modern science? Ancient Greece 3. Which philosopher originally thought up the atomic theory? Democritus 4.

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Unit 3 Atomic Structure- Guided Notes

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The Periodic Table ... □ Add you answers from number 1 together □ Round your answer to 2 decimal places □ Practice: 1. A sample of Cesium is 75% ^{133}Cs , 20% ^{132}Cs , and 5% ^{134}Cs . What is the average atomic mass? 2. Carbon-12 is 98% abundant and Carbon-14 is 2% abundant.

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Name: Period: Unit 3 Atomic Structure- Guided Notes

- Write the atomic mass at the bottom of the square. Step 2: Determine the number of protons, neutrons, and electrons in each element. Step 3: Create an atomic structure diagram for each element. Step 4: Draw

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the Bonding Structure for each element. 5
B Boron 10.81 P = 5 N = 6 E = 5 Atomic
Structure: Bonding Structure: B 5 P+ 6 no

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STRUCTURE DIAGRAMS Step 1:

Complete the squares for each element by
adding the atomic number, name and

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atomic mass. - Write the atomic number at the top of the square. - Write the element's name under the symbol. - Write the atomic mass at the bottom of the square.

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